A hot syngas purification system integrated with downdraft gasification of municipal solid waste

Wei Ping Chan⁎, Andrei Vekshaa, Junxi Leia, Wen-Da Oha,b, Xiaomin Doua, Apostolos Giannisac, Grzegorz Lisaka,d, Teik-Thye Limabcd,⁎

⁎ Corresponding authors at: Residues and Resource Reclamation Centre, Nanyang Environment and Water Research Institute, Nanyang Technological University, 1 Cleantech Loop, Clean Tech One, 637141, Singapore.
E-mail addresses: wpchan@ntu.edu.sg (W.P. Chan), cttlim@ntu.edu.sg (T.-T. Lim).

HIGHLIGHTS
• Demonstrated feasibility of MSW gasification with a hot syngas purification system.
• Produced hot and clean syngas through downward cascading of system temperatures.
• Removed up to 90% of tar compounds and sulfur species (H2S and COS).
• Demonstrated operational stability of Ni/Al2O3 catalyst for over 20h on stream.
• Demonstrated reliable performance of NiZn-28-HC sorbent with real MSW syngas.

ARTICLE INFO

Keywords:
Syngas purification
Municipal solid waste
Catalytic tar reforming
Nickel catalyst
Ni-Zn nanocomposite
Sulfur removal

ABSTRACT

Gasification of municipal solid waste (MSW) with subsequent utilization of syngas in gas engines/turbines and solid oxide fuel cells can substantially increase the power generation of waste-to-energy facilities and optimize the utilization of wastes as a sustainable energy resources. However, purification of syngas to remove multiple impurities such as particulates, tar, HCl, alkaline chlorides and sulfur species is required. This study investigates the feasibility of high temperature purification of syngas from MSW gasification with the focus on catalytic tar reforming and desulfurization. Syngas produced from a downdraft fixed-bed gasifier is purified by a multi-stage system. The system comprises of a fluidized-bed catalytic tar reformer, a filter for particulates and a fixed-bed reactor for dechlorination and then desulfurization with overall downward cascading of the operating temperatures throughout the system. Novel nano-structured nickel catalyst supported on alumina and regenerable Ni-Zn desulfurization sorbent loaded on honeycomb are synthesized. Complementary sampling and analysis methods are applied to quantify the impurities and determine their distribution at different stages. Experimental and thermodynamic modeling results are compared to determine the kinetic constraints in the integrated system. The hot purification system demonstrates up to 90% of tar and sulfur removal efficiency, increased total syngas yield (14%) and improved cold gas efficiency (12%). The treated syngas is potentially applicable in gas engines/turbines and solid oxide fuel cells based on the dew points and concentration limits of the remaining tar compounds. Reforming of raw syngas by nickel catalyst for over 20h on stream shows strong resistance to deactivation. Desulfurization of syngas from MSW gasification containing significantly higher proportion of carbonyl sulfide than hydrogen sulfide, traces of tar and hydrogen chloride demonstrates high performance of Ni-Zn sorbents.
1. Introduction

Global municipal solid waste (MSW) generation rate is estimated at 1.3 billion tonnes per year in 2012, and it is expected to increase to 2.2 billion tonnes annually by 2025 [1]. Incineration is commonly applied in the management of MSW to achieve 70–80% mass reduction and 80–90% volume reduction [2] while it produces electricity with up to 30% of net electrical efficiency [3]. Recently, development of different Waste-to-Energy (WtE) treatment technologies such as pyrolysis, gasification and hydrothermal processes [4] have attracted significant research interest as alternative to incineration [5]. Gasification is an advanced thermochemical conversion process which converts MSW into gaseous fuel (syngas) that can be utilized in gas engines, combined cycle gas turbines (CCGTs) [6], solid oxide fuel cells (SOFCs) [7], or hybrid systems [8] with higher electrical efficiency and better environmental performance [9] and therefore optimize the utilization of wastes as a sustainable energy resources. Formation of dioxins and furans is an important concern when MSW, particularly plastic wastes, are treated in thermochemical conversion processes. Nevertheless, as compared to incineration, reducing environment in a gasifier strongly inhibits the formation of these compounds [2]. In addition, Cl containing species (HCl, KCl and NaCl) can be removed from the syngas through the purification system, which further reduce the potential of dioxins and furans formation during the downstream applications of syngas [10]. Combustion of treated syngas as a homogenous gaseous fuel in gas engines/turbines instead of the direct combustion of highly heterogeneous MSW also reduce the formation of dioxins and furans. This is because the well-controlled complete combustion eliminates the residual carbon in the flue gas that acts as a precursor for the formation of dioxins and furans [11].

However, the main constraint of using syngas for downstream applications is the presence of high concentrations of impurities such as particulates, tar, HCl, alkali chlorides and sulfur species [12]. There are conventional and highly efficient cold clean-up systems such as venturi scrubbers, wash towers, wet/dry electrostatic precipitators, adsorbing beds or cyclones. These systems suffer from significant heat and energy efficiency losses and produce significant amount of solid and liquid waste streams [13]. Therefore, warm or hot clean-up systems are preferable to maintain high thermal efficiency of the WtE facilities and convert the impurities into potentially useful products [14].

Various warm/hot clean-up processes have been developed to remove different impurities in syngas. Particulates emitted from a gasifier consists of residual solid carbon and inorganic compounds such as alkali and alkaline earth metals, silica and trace constituents of other metals [15]. Various techniques are readily available for particulates removal such as cyclones, dust agglomerators, fabric filters, electrostatic separators and ceramic candles with removal efficiency of 99.5% or more at warm/hot temperature of 400 °C and above [16]. Chlorine species in syngas are predominantly represented as HCl, NH4Cl and alkali chlorides which are either vapors or solids depending on the operating temperatures [13]. Sodium- and calcium minerals can be applied as sorbents for removing HCl while activated alumina and natural minerals such as kaolinite and bauxite can be used as sorbents for removing alkali chlorides at around 400 °C or higher temperature [15].

Tar is a complex mixture of organic compounds with molecular weight greater than that of benzene. It consists of oxygenated products, heavier deoxygenated hydrocarbons and polycyclic aromatic hydrocarbons (PAHs) [17]. Tar compounds can be categorized based on their chemical structure, number of carbon ring and molecular weight [18]. The categorization and identification of tar components are crucial because the concentration limits [19] and dew points [20], the two most important criteria to be considered for downstream application of syngas, are significantly affected by the specific tar compounds found in the syngas. In general, a highly efficient syngas purification system should focus on the reduction of heavier tar compounds if complete removal of tar is not cost-effective [20].

Thermal cracking at extreme temperature of 1075–1300 °C, catalytic reforming with different types of catalysts and plasma treatment are typical hot/warm tar removal technologies [15]. Catalytic tar reforming is considered as one of the most promising technologies, considering its scalability and energy efficiency, but the main limitation is catalyst deactivation [21]. In addition, catalytic tar reforming with nickel-based catalysts showed superior tar reduction performances at high temperature [22], improved heating values of syngas, increased gas yield and demonstrated catalytic conversion of NH3 into N2 [23]. Existing studies focus extensively on fixed bed reactor with catalysts in the shape of pellet, ring, spherical, monolith and catalytic filter [21]. Nevertheless, catalytic reforming with a fluidized bed could improve the efficiency of this process by promoting the mixing of syngas with catalyst, maintaining uniform temperature in the tar reformer and reducing potential bed agglomeration and coking [24]. Nickel particles can be supported by using different materials which include alumina [25], limestone [22], dolomite [26] and char [27] to improve the catalytic activity and maintain the stability during reforming. Alumina can be considered as a suitable support material when high temperature fluidized-bed catalytic tar reformer is applied because of its high mechanical integrity, thermal stability [25], inert to gasification reactions and strong resistance to sintering and poisoning by HCl [22].

Sulfur species in syngas are predominantly present as H2S and COS [28]. Adsorption techniques that include chemical sorbents are commonly applied to remove sulfur. Metal oxides are typically used for high temperature desulfurization, while improvement on regenerability and removal efficiency of the sorbent materials are currently under development [15]. Advancement of hierarchical [29] and regenerable zinc oxide sorbents for syngas desulfurization suggested the potential of utilizing nanostructured materials to improve the techno-economic performance of desulfurization system [30]. Incorporation of additives (Ni, Cu, etc.) as promoters into ZnO sorbents could further enhance the desulfurization performance, increase the sorption capacity and stability, and improve the regenerability of the sorbents [31]. To further improve the applicability of sorbent, ZnO could be immobilized on a honeycomb which would minimize pressure drop and avoid pore plugging in the desulfurization reactor [28]. Regeneration of the used sorbents produces SO2 which can be subsequently recovered as sulfuric acid [32] or elemental sulfur [33].
From the thermal efficiency perspective, positioning of catalytic tar reformer at the outlet of gasifier followed by the downward cascading of syngas temperature through the purification system is highly efficient. This arrangement reduces the need for syngas cooling and re-heating, thus decreasing the cost and increasing the efficiency of purification system [34]. However, tar reforming catalysts could lose their reforming activity when directly in contact with raw syngas due to carbon and particulates deposition [35], poisoning by sulfur species, deactivation by HCl [36] and sintering [37]. On the other hand, desulfurization sorbents are difficult to be regenerated with highly consistent performance and sorbent deactivation can be accelerated due to the presence of chlorine species in syngas [29]. Our recent study has reported the synthesis of Ni/Al2O3 catalyst with high stability to poisoning by HCl due to the strong Ni nanoparticle-support interactions and nano-sized porous alumina support [22]. Additionally, Ni-Zn oxide nanocomposites have been developed with good regenerability and high sulfur capacity compared to pure ZnO in model syngas [28]. However, these advanced materials have not been used previously for the purification of real MSW syngas and there is limited research in the arrangement of multi-stage system with downward cascading of temperature for the purification of real syngas produced from MSW gasifier. Therefore, testing of these novel catalysts and sorbents in real syngas with wide range of impurities through a well arranged multi-stage system is of great importance to understand the system stability and efficiency throughout the purification processes.

In this study, a novel syngas purification system is constructed to investigate the feasibility of high temperature purification of raw syngas from gasification of MSW, with the focus on catalytic tar reforming, desulfurization and the system arrangement with downward cascading of syngas temperatures. MSW, in the form of refuse-derived fuel (RDF) pellets, is gasified in a downdraft fixed-bed gasifier at 850 °C with equivalence air ratio (ER) of 0.3 and 25wt% moisture content to produce syngas with consistent quality similar to the properties of syngas generated in industrial gasifiers. The generated raw syngas is then treated in a multi-stage purification system with downward cascading of the syngas temperature to remove impurities while producing hot and clean syngas at 400 °C as a final product. For the first time, a fluidized-bed tar reforming reactor with Ni catalysts supported on Al2O3 is utilized for the removal of tar from real MSW syngas containing particulates, chlorine and sulfur compounds. Particulate filter and Na2CO3 are installed to remove particulates and HCl prior to the desulfurization unit. The removal of sulfur species (H2S and COS) is reported in a previous study [28]. The removal of sulfur species (H2S and COS) is carried out by using Ni-Zn oxide sorbent. Efficiency of the removal of impurities, stability of catalysts and sorbents, and performance of the purification system are analyzed. Selectivity of tar reforming catalysts towards tar decomposition from MSW syngas is described. Dew points and specific contents of different tar compounds in the treated syngas are determined and compared to the operating requirements and concentration limits of the downstream application of syngas, which include gas engines/turbines and SOFCs. Complementary sampling and analysis methods are applied for tar, alkali chlorides and HCl quantification to provide insights into the distribution of these impurities throughout the multi-stage system. Experimental and thermodynamic modeling results are compared to determine the differences caused by the kinetic constraints of the multi-stage system.

2. Materials and methods

2.1. Preparation and characterization of RDF

MSW generated in Nanyang Technological University (Singapore) was used to prepare the RDF pellets. Approximately 200 kg of mixed wastes were collected from the central waste collection site. The wastes were sorted, weighed and dried according to ASTM D5231. Six main combustible components, plastic, paper, textile, wood, food residues and horticultural wastes were selected and retained to form the RDF pellets while other waste components were discarded, which include glassware, metals, liquids, stones, ceramics and rubbers. The overall moisture content of the collected wastes was 25 wt%. After drying at 105 °C for at least 24 h, the components were crushed into pieces (Retsch SM2000 and plastic crusher, DJ400). Crushed wastes were mixed and homogenized to form a mixture with 35 wt% plastic, 25 wt% paper, 7.7 wt% textile, 6.9 wt% wood, 16 wt% food residues and 11.4 wt% horticultural wastes. Homogenized mixture was pelletized to form RDF pellets with diameter of 4 ± 1 mm and length of 10 ± 3 mm (pellet mill, Gemco). Characteristics of RDF pellets (Table S1) were determined based on the methods described previously [38].

2.2. Catalysts, filter and sorbent materials

One commercially available catalyst (Pingxiang Hualian Ceramic Co., China) and one fabricated nickel catalyst supported on alumina were used for fluidized-bed tar reforming. For the catalyst fabrication, aluminum hydroxide H3AlO3·xH2O (Sigma-Aldrich) was pelletized and the prepared pellets were crushed to obtain particles with sizes 0.56–1.18 mm. Approximately 17.2 g of particles were added to the solution containing a known quantity of Ni(NO3)2·6H2O (Sigma-Aldrich) in 10 mL deionized water and mixed. The solvent was evaporated using a rotary evaporator Hei-Vap Precision (Heidolph Instruments). The material was dried overnight in an oven at 105 °C and then calcined in air at 850 °C for 2 h (heating rate of 2 °C/min). Both commercial and fabricated catalysts were crushed and sieved to the particle size of 100 to 315 μm and are denoted as C-Cat and F-Cat, respectively. Further details on chemicals, synthesis procedure and characterization of these catalysts could be found elsewhere [22]. Bulk densities of C-Cat and F-Cat were 1.27 g/mL and 0.80 g/mL respectively. Based on the X-ray fluorescence (XRF, PANalytical Axios mAx), NiO contents for C-Cat and F-Cat were 19.9 wt% and 15.8 wt% respectively. 3 mL of the C-Cat (space velocity, SV = 6000 h⁻¹) and 3 mL, 4 mL and 5 mL of F-Cat (SV = 6000, 4500 and 3600 h⁻¹, respectively) were used in the fluidized-bed tar reformer. The effective loadings of NiO calculated from XRF data were 0.76 g for 3 mL of C-Cat and 0.38, 0.51 and 0.63 g for 3, 4 and 5 mL of F-Cat, respectively.

A quartz thimble filter (No.88R, Advantec) was used to remove particulates. CaO and Na2CO3 (Sigma-Aldrich) were pelletized, crushed and sieved to a particle size of 2.0–2.8 mm and used for the HCl capture.

Ni-Zn nanocomposite loaded on honeycomb (cordierite mullite), denoted as NiZn-28-HC, was applied to remove sulfur species (H2S and COS). For the preparation of NiZn-28-HC, the honeycomb was first seeded by immersing into a seeding solution containing 20 mmol of Zn (CH3COO)2·2H2O in 50 mL ethanol and 2 mL diethanolamine for 2 h [30]. The honeycomb was then slowly removed from the seeding solution, dried at 60 °C for 2 h and calcined in air at 400 °C for 2 h (heating rate of 2 °C/min). The seeded honeycomb was placed into a solution prepared from 5 mmol of Zn(NO3)2·6H2O, 20 mmol of urea and 2 mmol of Ni(NO3)2·6H2O in 30 mL deionized water under rapid magnetic stirring for 20 min, followed by hydrothermal synthesis at 100 °C for 18 h and calcination at 400 °C for 5 h (heating rate of 2 °C/min). Details of the synthesis and characterization of this sorbent are described in a previous study [28]. Mass of the honeycomb was 3.06 ± 0.05 g (SV = 11,500 h⁻¹) with 1.77 ± 0.02 wt% of Ni-Zn and Ni:Zn molar ratio of 2.5 on the honeycomb.

2.3. Characterization of materials

Temperature programmed reduction (TPR) was performed by flowing a 5% H2/N2 gas mixture at 30 mL/min and a heating rate of 10 °C/min up to 900 °C. Transmission electron and field-emission scanning electron micrographs were obtained by using JEOL microscopes (JEM-1400 and JEM 2010 for TEM and 7600F for FESEM). X-ray diffraction (XRD) patterns were obtained by using X-ray diffractometer.
(Bruker AXS D8 Advance) operated with high intensity monochromatic Cu-Kα source at λ = 1.5418 Å at 40 kV and 40 mA. X-ray photoelectron spectroscopy (XPS) was conducted by using a spectro-photometer (Kratos Axis Supra) with a dual anode monochromatic Kα excitation source. All binding energies for elements of interest were corrected against an adventitious carbon C 1s core level at 284.8 eV. XPS peaks were deconvoluted using CASA XPS software by fitting the peaks on Shirley background coupled with Gaussian-Lorentzian function.

2.4. Gasification and syngas purification system

The multi-stage purification system integrated with gasifier is illustrated in Fig. 1. A downdraft fixed-bed gasifier was used to produce raw syngas by gasifying RDF pellets with a feeding rate of approximately 4.4 ± 0.3 g/min. Stoichiometric air-to-fuel (A/F) ratio was estimated as 6.96 L/g, by calculating stoichiometric balance of the complete oxidative reaction between C, H, N, S and O elements in RDF pellets and O2 in air assuming the reaction products are CO2, H2O, NO2 and SO2. To maintain an equivalence air ratio of 0.3 and water content of 25 wt%, 9.2 L/min of air and 1.47 mL/min of deionized water were injected simultaneously into the gasifier. Approximately 12 L/min of raw syngas (on a dry basis) was generated from the gasifier and 300 mL/min of the raw syngas was directed into the purification system. The system included a fluidized-bed catalytic tar reformer, a particulate filter and a fixed-bed reactor with dechlorination and desulfurization adsorbents to remove tar, particulates, HCl and sulfur species (Fig. 2) from the raw syngas with downward cascading of the syngas temperature. Connecting lines and particulate filters were maintained at 300 °C. Pressure in the gasifier and purification system were monitored continuously throughout the experiment by using pressure gauge (Omega) and digital manometer (Dwyer).

A typical experimental run was initiated with the pre-heating of gasifier, syngas purification system and all connecting lines to the designated operating temperatures. During heating, N2 (300 mL/min) was purged into the gasifier to maintain an inert environment while a mixture of H2 (50 mL/min) and N2 (50 mL/min) was purged into the syngas purification system to maintain a reducing environment for the reduction of tar reforming catalyst and desulfurization sorbent.

Approximately 15 min before the start of experiment, purging of H2 ceased while purging of N2 continued to remove the H2 and maintain an inert environment in the system. After pre-heating and stabilization of the temperature in gasifier, feeding of RDF pellets was started by using a screw-feeder which consisted of a fuel holder (mass capacity of 1.6 kg and volume capacity of 6 L) and a horizontal screw-conveyor (length of 380 mm and diameter of 50 mm). As the screw-conveyor turned, RDF pellets were crushed, compacted and pushed to the top of the gasifier. RDF then dropped from the top towards the bottom of the gasifier and onto a stainless-steel mesh which acted as a support for the formation of char bed. Simultaneously, air was fed by using a mass flow controller (FMA5500A, Omega) through a circular distributor near the central region of the gasifier, positioned at ~10–15 cm above the char bed and water was injected using a peristaltic pump (Masterflex L/S®, Cole Palmer) through a coiled stainless-steel pipeline. Water was heated along the pipeline and evaporated into steam when entering the gasifier. Two thermocouples (TC1 and TC2) were placed at the center of gasifier and char bed, respectively, to measure real-time temperature profiles inside the gasifier. The height of the gasifier was 585 mm and the char bed was positioned at 105 mm from the bottom of the gasifier. After the stabilization of syngas composition, raw syngas was directed into the purification system.

2.5. Sampling and analysis

During the experiments, sampling of syngas components was carried out at the outlets of the downdraft gasifier (GS), tar reformer (TR) and desulfurization reactor (DES). Syngas, tar, chlorides (e.g. NaCl and KCl), HCl, sulfur species, H2O, particulates were collected, analyzed and quantified. All sampling pumps (GilAir Plus, Sensidyne) and mass flow controllers were calibrated by using a bubble flow meter and checked during the experiments. Results were reported as averages ± standard deviations (SDs) of three runs. 2-sample t-test was carried out using Minitab® 17 to determine the statistical significance of the experimental results.

Sampling of tar was carried out using two different methods for comparison and complementary data collection. Method I is a modified method with a series of sampling traps using isopropanol (IPA) as
solvent based on previous report [39]. The sampling train included a quartz thimble filter, connecting lines, five impinger bottles filled with 50 mL of IPA each and one empty impinger bottle. The second, fourth, fifth and the sixth impinger bottles were equipped with fine-meshed frits to produce smaller gas bubbles and to improve efficiency of tar capture by IPA. All the parts were rinsed with IPA after the sampling. Soxhlet extraction was used to dissolve tar compounds accumulated on the thimble filter with 250 mL IPA. All the collected IPA solutions were homogenized in a volumetric flask, topped up to 1 L, and denoted as “collected tar”. Method II utilized solid phase adsorption (SPA) by using tubes loaded with 500 mg of aminopropyl-bonded silica gel (Bond Elut NH2, Agilent) [40]. Approximately 150 mL of syngas was directed through a SPA tube by using a sampling pump. Tar sampling was carried out at 30 min intervals simultaneously before and after the catalytic tar reformer (GS and TR sampling points, respectively). After sample collection, SPA tubes were sealed and stored in a dark cold room at 4°C for 24 h before elution and analysis. A mixture of dichloromethane, acetonitrile and IPA with a volume ratio of 8:1:1 was used to extract polar and non-polar compounds simultaneously from the SPA tubes [41]. The tar analyzed by Method II is denoted as “SPA tar”.

Collected tar was sampled continuously throughout the experiment while SPA tar was sampled intermittently at selected time points. Therefore, collected tar represents the cumulative tar content while SPA tar represents instantaneous tar contents at different sampling times. Since the volume of IPA for Method I was 1 L, a large volume of syngas was needed to pass through the series of sampling traps to accumulate sufficient amount of tar for detection and quantification. In contrast, by Method II, small amount of syngas was required as only 2 mL of solvent was needed for tar extraction from SPA tubes. Method II is therefore simpler and faster as compared to Method I. In addition, SPA tar collection could detect the evolution of tar in the syngas during the experiments and occasional escape of tar compounds from the catalytic tar reformer.

Quantification of tar compounds in collected tar and SPA tar was carried out by using GC–MS (HP7890 GC with a 5975 MS, Agilent). The gas chromatograph was equipped with a HP-INNOWAX capillary column (i.d. of 0.250 mm and a film thickness of 0.25 µm). Chromatograms were recorded in total ion current (TIC) mode and peaks were identified by comparing their mass spectra to the NIST 11 mass spectra library. 33 tar compounds were quantitatively analyzed. The quantified tar compounds are shown in Table S2. Deuterated d8-toluene and d8-naphthalene were applied as internal standards for the analysis of all samples. The sum of concentrations of all tar compounds detected by GC–MS is denoted as total tar. Tar dew point, defined as the temperature at which the real total partial pressure of tar and the saturation pressure of tar are equal, was calculated based on the model developed by Energieonderzoek Centrum Nederland (ECN) and the concentrations of 34 individual tar components [42]. Tar removal (%) was calculated from the tar contents in raw and treated syngas (Tarraw and Tarcreated) respectively:

\[
\text{Tarremoval} \% = \frac{\text{Tar}_\text{raw} - \text{Tar}_\text{created}}{\text{Tar}_\text{raw}} \times 100\% \quad (1)
\]

To analyze HCl and the main gases, syngas was pumped through two impinger bottles with deionized water to absorb HCl and then passed through a trap with silica gel to remove moisture and collect dry gas into a gas bag. Concentration of HCl in syngas was calculated based on the content of chloride anion (mg/L) in water traps analyzed by ion chromatograph (ICS-1100, Dionex). The amount of Cl− associated with K+ and Na+ was measured using ICP-OES (Optima 8300, Perkin Elmer). The amount of Cl− associated with K+ and Na+ was calculated and deducted from total Cl− measured to quantify the concentration of HCl. The collected syngas was analyzed by a GC (7890B GC system, Agilent) coupled with two thermal conductivity and one flame ionization detectors. Details of quantified gas components are shown in Table S4.

Volume of the generated syngas is estimated based on the fraction of N2 in air and syngas, respectively, and assumed that of N2 is inert during the gasification and syngas purification. Output of syngas was then calculated as follows:

\[
\text{Nm}_\text{input} = \text{Air}_\text{input} \times \text{Fraction of N2 in air} \quad (2)
\]

\[
\text{Syngas}_\text{output} = \text{Nm}_\text{input} \times \text{Fraction of N2 in syngas} \quad (3)
\]

where Syngasoutput is the output of syngas on a dry basis, Airinput is the input of air and Nminput is the amount of N2 in air.

Lower heating value (LHV) and cold gas efficiency (CGE) of syngas were calculated based on the heating values of the components in syngas (Table S5):

\[
\text{CGE} = \frac{\text{LHV}_\text{output}}{\text{LHV}_\text{input}} \times 100\% \quad (4)
\]

\[
\text{LHV}_\text{output}(\text{MJ}) = \text{LHV}_\text{syngas}(\text{MJ}/\text{Nm}^3) \times \text{Syngas}_\text{output}(\text{Nm}^3) \quad (5)
\]

\[
\text{LHV}_\text{input}(\text{MJ}) = \text{LHV}_\text{sys}(\text{MJ}/\text{kg}) \times \text{Mass}_\text{sys}(\text{kg}) \quad (6)
\]

where LHVoutput is the total energy output in the syngas and LHVinput is the total energy input into the gasifier through RDF.

For the quantification of sulfur species (H2S and COS), the collected syngas was analyzed by a GC (7890B, Agilent) equipped with a flame photometric detector. The content of particulates was calculated from the mass of quartz thimble filter (dried at 105°C for 24 h) before and after each experiment. Solid residues remained in the gasifier after the experiments including char on the char bed, ashes in the ash tray and deposits on the inner wall of gasifier were collected, weighed and characterized.

2.6 Thermodynamic modeling

Thermodynamic modeling was carried out by using HCS Chemistry 9.0 (GEM, Outotec) which applied Gibbs energy minimization method to calculate the contents of CO, CO2, CH4, H2, C2-C5 gases and H2O at the equilibrium. The elemental composition of RDF pellets, amount of air input (ER = 0.3), water input (25 wt%), temperature (850°C) and pressure (101.3 kPa) were used as the input parameters.

Fig. 2. Flow diagram of the hot syngas purification system integrated with downdraft gasification of municipal solid waste. Materials, operation temperatures and impurities removed are indicated.
3. Results and discussion

3.1. RDF gasification and properties of raw syngas

Temperature profiles of the oxidation (near air distributor) and reduction (char bed on the stainless-steel mesh) regions as denoted by TC1 and TC2 are illustrated in Fig. S1. Based on the smaller standard deviations, relatively consistent temperatures were observed in the oxidation zone (883 ± 8°C) compared to the reduction zone (873 ± 22°C). The production and composition of raw and treated syngas were generally stable as illustrated in Fig. 3a for raw syngas and Fig. 3b for treated syngas. The syngas generated from the gasifier, consisted of 13 ± 2 vol% CO, 11 ± 0.4 vol% CO2, 3.2 ± 0.2 vol% CH4, 12 ± 2 vol% H2, 2.9 ± 0.3 vol% C2-C5 gases and 58 ± 4 vol % N2 (Fig. 4a), with a syngas yield (dry basis) of 2.7 ± 0.3 and 1.1 ± 0.2 Nm³/kg of RDF with and without N2, respectively. LHV of the raw syngas was 6.7 ± 1.6 MJ/Nm³ while the CGE of this althermal gasifier was 79 ± 20%. Moisture content (wet basis) in syngas generated from gasifier was 23 ± 4 vol% while the particulates content was 1.9 ± 1.2 g/Nm³. Collected and SPA tar contents (Table 1) in raw syngas were 7.8 ± 3.7 g/Nm³ and 7.3 ± 1.6 g/Nm³, respectively, suggesting no statistically significant difference between the two analytical methods. 25 different tar compounds (Table S3) were similarly identified in the collected and SPA tar from raw syngas. Toluene, styrene and naphthalene were the main tar compounds which accounted for 74 ± 3 wt% of total tar content. The contents of styrene in collected and SPA tar were 1.4 ± 0.8 g/Nm³ and 1.7 ± 0.5 g/Nm³, respectively. This is different from the syngas generated during biomass gasification, in which styrene is a minor tar component [13]. Formation of styrene could be attributed to the presence of polystyrene in the MSW [43].

The content of solid residues, which were mixtures of char from the char bed, ashes from the ash tray and deposits on the inner wall of the gasifier, was 15 ± 4 wt% (based on the total amount of RDF input). Characterization results showed that the solid residues consisted of 26 ± 4 wt% of C, 0.7 ± 0.1 wt% of H, 0.5 ± 0.1 wt% of N, 0.4 ± 0.1 wt% of S, 5 ± 3 wt% of O (calculated by difference), 68 ± 4 wt% of ashes, and 4 ± 0.5 wt% Cl. Syngas, tar and solid residues contained 84 ± 12%, 6 ± 1% and 4 ± 2% of total carbon, respectively, and carbon balance closure was 94 ± 12%.

After the correction for chlorine content contributed by the alkali chlorides, the HCl content in syngas was 0.26 ± 0.15 g/Nm³ as showed in Table 3. Concentrations of KCl and NaCl were 0.44 ± 0.17 g/Nm³ and 0.56 ± 0.23 g/Nm³ respectively. Mass ratio for K:Na was approximately 1:1. HCl, KCl, and NaCl in raw syngas accounted for 9 ± 5%, 7 ± 3% and 11 ± 5% of total Cl, respectively, while the solid residues contained for 55 ± 16% of total Cl. Therefore, 82 ± 18% closure of Cl mass balance was achieved. The losses of Cl could be attributed to the condensation of alkali chlorides in the pipes and reactors.

Approximately 80 ppmv of sulfur species (H2S and COS), as showed in Table 2, were detected in the syngas generated from the gasifier. Evolution profiles of H2S and COS during a typical experiment carried out in the experimental set-up are presented in Fig. S2. Concentration of COS in raw syngas was higher compared to that of commonly reported for biomass and coal syngas, in which H2S is the main sulfur compound [44], which can be attributed to the differences in the characteristics of feedstock and produced syngas [45]. Nevertheless, comparable poisoning effects on the Ni catalysts by H2S and COS were previously reported [46]. Therefore, the Ni catalysts treating raw syngas from MSW gasification were required to be strongly resistant to both sulfur species and in this case, the total amount of sulfur species should be considered as a crucial parameter, equally important as the concentrations of individual sulfur compounds in the syngas.

3.2. Fluidized bed tar reforming

Since the temperature of syngas at the outlet of gasifiers is typically high, it is energetically favorable to carry out tar reforming without cooling down of syngas. In addition, it is important to carry out catalytic tar reforming at temperature higher than 800°C to suppress the poisoning effect of sulfur species on the Ni catalysts [47]. Although the removal of particulates prior to tar reformer would be beneficial [48], filter materials that can operate at such high temperatures are still under development [49]. Therefore, a fluidized-bed catalytic tar reformer was applied in the purification system to avoid plugging caused by particulates in raw syngas. Pressure of the purification system was constant throughout the experiment suggesting no significant bed agglomeration and plugging when raw syngas was directed into tar reformer without any pre-treatment.

Experiments with an empty reforming reactor (empty reformer) showed that the total tar content (Table 1), composition of main gases (Fig. 4), gas yield and LHV (Fig. 5) of syngas were statistically similar to those of raw syngas. However, the number of detected tar compounds was reduced to 15 (Table S3) based on the results from collected tar samples. Styrene and some other tar compounds were partially decomposed, while the concentration of naphthalene significantly increased, indicating that thermal treatment changed the composition of tar due to thermal cracking and conversion of tar compounds to naphthalene. This is probably because of the high thermal stability of naphthalene as compared to other tar compounds [50]. In the presence of catalysts in the tar reformer, most of the tar compounds were reformed and the total tar contents in collected tar decreased to 0.8 ± 0.3 g/Nm³ with 5 mL of F-Cat, suggesting up to 90% of tar removal efficiency (Table 1). The only tar compounds that were detected in the treated syngas were toluene, styrene and naphthalene, with toluene and naphthalene accounting for more than 95 wt% of the collected tar. These results suggest that further research should be carried...
out with the focus on the design of catalysts with high selectivity to the reforming of these two tar compounds. Fig. 6a presents the removal of collected tar and individual tar components using the F-Cat and C-Cat. The removal efficiency of the catalysts is plotted against NiO loading in the catalyst bed. Removal of total tar, toluene, naphthalene and styrene increased with the increased NiO loading from 0.38 to 0.51 g. The further increase in NiO loading did not decrease tar content significantly suggesting that the tar removal efficiency reached a plateau and could not be significantly improved with only the increment of NiO loading. Fig. 6b and 6c illustrate collected and SPA tar removal using 4 mL F-Cat (NiO loading of 0.51 g) for syngas reforming. F-Cat demonstrated operational stability in direct contact with the raw syngas for more than 20 h on stream (Fig. 6b). The fluctuations in removal efficiencies during gasification runs could be observed using SPA tar samples (Fig. 6c). These fluctuations could be one of the reasons that reduced the overall tar removal efficiency.

It is important to assess the potential of using the treated syngas in real applications. The assessment was carried out based on the dew points and concentration limits of tar. Tar dew point is an important parameter to be considered for downstream applications of syngas [51]. Typically, the gas intake system of gas engine/turbine has to be maintained at approximately 28 °C above the dew point of the syngas to avoid any condensation of tar [52]. The calculated tar dew point of the treated syngas was 22 ± 6 °C. Therefore, by maintaining the operating temperature of the gas intake system at around 50 °C, the treated syngas is potentially usable in the gas engine/turbine. On the other hand, concentration limits set based on CEN/BT/TF 143 for different categories of tar compounds according to the number of carbon ring and

Table 1

<table>
<thead>
<tr>
<th>Sampling method</th>
<th>Test</th>
<th>Total Tar (g/Nm³)</th>
<th>Toluene</th>
<th>Styrene</th>
<th>Naphthalene</th>
</tr>
</thead>
<tbody>
<tr>
<td>Collected tar</td>
<td>Gasifier</td>
<td>7.8 ± 3.7</td>
<td>2.7 ± 1.5</td>
<td>1.4 ± 0.8</td>
<td>1.7 ± 0.6</td>
</tr>
<tr>
<td></td>
<td>Empty reformer</td>
<td>7.2 ± 1.9</td>
<td>2.8 ± 0.5</td>
<td>0.9 ± 0.2</td>
<td>3.0 ± 0.9</td>
</tr>
<tr>
<td></td>
<td>C-Cat 3 mL</td>
<td>1.1 ± 0.5</td>
<td>0.6 ± 0.3</td>
<td>0.1 ± 0.1</td>
<td>0.5 ± 0.3</td>
</tr>
<tr>
<td></td>
<td>F-Cat 3 mL</td>
<td>2.8 ± 0.2</td>
<td>1.6 ± 0.2</td>
<td>0.2 ± 0.03</td>
<td>0.9 ± 0.04</td>
</tr>
<tr>
<td></td>
<td>F-Cat 4 mL</td>
<td>1.2 ± 0.5</td>
<td>0.6 ± 0.2</td>
<td>0.02 ± 0.03</td>
<td>0.5 ± 0.3</td>
</tr>
<tr>
<td></td>
<td>F-Cat 5 mL</td>
<td>0.8 ± 0.3</td>
<td>0.5 ± 0.1</td>
<td>0.04 ± 0.06</td>
<td>0.3 ± 0.1</td>
</tr>
<tr>
<td>SPA tar</td>
<td>Gasifier</td>
<td>7.3 ± 1.6</td>
<td>2.9 ± 0.9</td>
<td>1.7 ± 0.5</td>
<td>1.5 ± 0.2</td>
</tr>
<tr>
<td></td>
<td>C-Cat 3 mL</td>
<td>1.0 ± 0.6</td>
<td>0.2 ± 0.1</td>
<td>0.03 ± 0.02</td>
<td>0.7 ± 0.5</td>
</tr>
<tr>
<td></td>
<td>F-Cat 3 mL</td>
<td>1.1 ± 0.5</td>
<td>0.5 ± 0.3</td>
<td>0.1 ± 0.04</td>
<td>0.6 ± 0.4</td>
</tr>
<tr>
<td></td>
<td>F-Cat 4 mL</td>
<td>0.6 ± 0.3</td>
<td>0.2 ± 0.1</td>
<td>0.02 ± 0.02</td>
<td>0.4 ± 0.2</td>
</tr>
<tr>
<td></td>
<td>F-Cat 5 mL</td>
<td>0.7 ± 0.4</td>
<td>0.2 ± 0.03</td>
<td>0.03 ± 0.02</td>
<td>0.5 ± 0.3</td>
</tr>
<tr>
<td>Corrected SPA tar</td>
<td>C-Cat 3 mL</td>
<td>1.3 ± 0.8</td>
<td>0.5 ± 0.2</td>
<td>0.03 ± 0.02</td>
<td>0.7 ± 0.5</td>
</tr>
<tr>
<td></td>
<td>F-Cat 3 mL</td>
<td>2.1 ± 0.9</td>
<td>1.5 ± 0.6</td>
<td>0.1 ± 0.04</td>
<td>0.6 ± 0.4</td>
</tr>
<tr>
<td></td>
<td>F-Cat 4 mL</td>
<td>0.9 ± 0.5</td>
<td>0.5 ± 0.01</td>
<td>0.02 ± 0.02</td>
<td>0.4 ± 0.2</td>
</tr>
<tr>
<td></td>
<td>F-Cat 5 mL</td>
<td>0.9 ± 0.5</td>
<td>0.4 ± 0.03</td>
<td>0.03 ± 0.02</td>
<td>0.5 ± 0.3</td>
</tr>
</tbody>
</table>

![Fig. 4. Gas composition of (a) raw syngas generated from gasifier (GS) and (b) treated syngas after passing through tar reformer (TR) for five different sets of experiments, which included empty reformer (without catalyst) and catalytic tar reformer placed with C-Cat (3 mL) and F-Cat with different dosages (3 mL, 4 mL and 5 mL). The data are showed as averages ± SDs of three runs.](image)
molecular weight as showed in Table S6 suggested that the treated syngas could be applied in the gas engines/turbines [19]. The contents of one ring (toluene and styrene) and two rings (naphthalene) tar compounds in the treated syngas were 0.5 ± 0.2 and 0.3 ± 0.1 g/Nm³ respectively. These concentrations were lower than the equivalent limits of 9.5 and 1.3 g/Nm³ respectively. Alternatively, the treated syngas could be utilized in SOFC for electricity generation which demonstrated successful operation with the syngas consisted of 0.5–10 g/ Nm³ of tar [53].

Nevertheless, it is still desirable to further reduce the tar content in the syngas to achieve high performance in the integrated gasification, purification and syngas utilization system. The improvement could be carried out by further testing of different dosages of catalysts, operating conditions of catalytic tar reformer, synthesis of novel catalysts and the addition of polishing unit employing adsorption/absorption techniques with the use char [54], activated carbon and waste oil [55].

In contrast to tar sampled at the outlet of gasifier (~300 °C, 101 kPa), statistically significant differences were observed for tar sampled at the outlet of the tar reformer after catalytic treatment (~850 °C, 95 kPa) among the samples of collected and SPA tar. Differences between the data obtained from the two tar sampling methods for tar contents measured after catalytic tar reformer are illustrated in Fig. 7. Results showed that total tar contents estimated by using SPA were sometimes lower than the collected tar. Toluene contents were consistently lower in SPA than collected tar. Comparable styrene and naphthalene contents were observed for both methods. Correlation analysis revealed that the difference of total tar content estimated by these two methods could be primarily accounted by the difference between the toluene contents as illustrated in Fig. 8a, with a coefficient of determination (R²) of 0.88. The loss of toluene could be attributed to the higher temperature and lower pressure at the sampling point after tar reformer which promote devolatilisation of the volatile organic and light tar compounds [56]. Nevertheless, the underestimation or loss of toluene content by SPA sampling method was closely correlated with the toluene content in syngas measured by using collected tar, as illustrated in Fig. 8b with a strong R² of 0.94. Therefore, tar sampling by SPA, though underestimated the contents of light tar compounds, could be corrected based on the supplementary results from collected tar and the corresponding correlation equations.

The corrected results of SPA tar were comparable to collected tar as showed in Table 1. SPA can also be applied as a quick and simple tar measurement method for naphthalene or heavy tar compounds which can be used as representative indicators of tar contents in the syngas for an industrial environment. These results suggested that both methods are suitable for tar sampling. The selection of sampling method should then be subjected to the sampling need and the suitability of environment for setting up the sampling system.

Fig. 4b illustrates the compositions of main gases for catalytic tar reformer for different catalysts at different dosages. No statistically significant difference was observed in the composition of main gases, syngas yield and LHV when two catalysts at different loadings were used. In addition, the changes between compositions of raw and corresponding treated syngas were statistically comparable when two catalysts at different loadings were used. Therefore, the apparent differences among the experiments were due to the variations in the composition of raw syngas (Fig. 4a). To account for these variations and simplify the further discussion, an average of the composition of treated syngas from catalytic tar reformer with two catalysts at different dosages was calculated. In average, after catalytic tar reforming, syngas contained 20 ± 2 vol% CO, 7 ± 0.6 vol% CO₂, 2.2 ± 0.1 vol% CH₄, 20 ± 1 vol% H₂, 0.4 ± 0.1 vol% C₂-C₅ gases and 51 ± 2 vol% N₂. Moisture content in the treated syngas was reduced to 18 ± 5 vol%.

Significant changes on the composition of raw gases and moisture content suggest the presence of steam reforming reactions which convert tar compounds, CH₄ and C₂-C₅ gases into H₂ and CO. Concurrently, proportion of CO₂ was reduced (Fig. 4b) which could be attributed to the dry reforming reactions and the increased volume of the treated syngas (Fig. 5). Total volume of syngas (on a dry basis) increased by 14 ± 5% to 3.1 ± 0.4 Nm³/kg of RDF pellets. Average LHV of syngas increased by 14 ± 8% (Fig. 5). The estimated CGE of this system was 91 ± 18% after the allothermal catalytic tar reformer.

Based on thermodynamic modeling using Gibbs energy minimization method, the theoretical composition of syngas (on a dry basis) was 20.1 vol% CO, 7.4 vol% CO₂, 27.8 vol% H₂ and 44.7 vol% N₂. The contents of CH₄ and C₂-C₅ gases were negligible. The calculated LHV and yield of theoretical syngas were 5.5 MJ/Nm³ and 3.7 Nm³/kg respectively. Comparison of the properties of raw, treated and theoretical syngas at thermodynamic equilibrium showed that experimental and modeling results were different, suggesting the presence of kinetic constraints in the gasification and syngas purification system. Nevertheless, as compared to raw syngas, the composition of treated syngas was closely correlated with the toluene content in syngas measured by using collected tar, as illustrated in Fig. 8b with a strong R² of 0.94. Therefore, tar sampling by SPA, though underestimated the contents of light tar compounds, could be corrected based on the supplementary results from collected tar and the corresponding correlation equations.
Syngas was closer to the theoretical values (Table 4) suggesting that catalytic tar reforming helps in pushing the composition of syngas towards thermodynamic equilibrium. The main reason for the deviations between experimental results and calculations could be attributed to the kinetic constraints of H$_2$O related reactions because significantly higher amount of H$_2$O and lower amount of H$_2$ are present in the raw and treated syngas as compared to the theoretical values. Three potentially incomplete reactions were suggested which include steam reforming of tar, CH$_4$ and C$_2$-C$_5$ gases [57], water-gas shift reaction and gasification of particulates in syngas and solids residues remained in the gasifier. This is demonstrated by a three-step calculation of the yields (Nm$^3$/kg) for the main gases (Table 4) using treated syngas as the initial input. These calculation steps were carried out consecutively to reach the gas composition of theoretical syngas at equilibrium. Step 1, the yields were calculated by applying stoichiometric calculation and assuming complete steam reforming of the hydrocarbons (tar, CH$_4$ and C$_2$-C$_5$ gases) in the treated syngas. Step 2, using the calculated yields from Step 1 and thermodynamic modeling with Gibbs energy minimization method, conversion of H$_2$O and CO through water-gas shift reaction and gasification of particulates in syngas and solids residues remained in the gasifier. The final calculated values at Step 3 and modeling results showed good agreement. This observation supported the hypothesis on the kinetic constraints of H$_2$O related reactions and incomplete gasification of solid residues. The kinetic constraints could be attributed to the low reactivity of char gasification with CO$_2$ and H$_2$O [58] and the poisoning of the catalysts by HCl [22] and sulfur species [59].

Concentrations of particulates in the syngas after passing through the catalytic tar reformer with 3 mL of C-Cat and 4 mL of F-Cat were 0.7 ± 0.3 and 0.8 ± 0.8 g/Nm$^3$, respectively. Slight reduction of particulates in syngas after TR could be attributed to partial deposition of carbon and inorganic compounds over the catalysts and partial decomposition by gasification of residual solid carbon [15]. This observation is supported by the increase in sulfur species contents (Table 2), which probably originated from the decomposed particulates and the reduced content of chlorides (Table 3) in the treated syngas. Concentration of HCl after TR was 0.10 ± 0.06 g/Nm$^3$. KCl and NaCl contents after TR were reduced to 0.08 ± 0.08 g/Nm$^3$ and 0.10 ± 0.10 g/Nm$^3$ respectively. The decrease in HCl, NaCl and KCl contents could be attributed to the deposition or adsorption of chlorides as reported previously for Ni-based sorbent [34] and alumina composites [60].

Approximately 90 ppmv of sulfur species (H$_2$S and COS), as showed in Table 2, were detected in the syngas after the catalytic tar reformer with 4 mL of F-Cat. Higher concentration of COS than H$_2$S was observed (Fig. S2), which is consistent with the distribution of sulfur species in the raw syngas. The slight increase of the total amount of sulfur species could be attributed to the conversion of sulfur-containing char and particulates into gases and partial hydrolysis or hydrogenation of COS to H$_2$S.

Since raw syngas was reformed directly in the catalytic tar reformer
without any pre-treatment, structural changes and properties of catalysts before and after the syngas purification were investigated. XRD patterns of the pristine and spent catalysts are showed in Fig. 9a. NiO diffraction peaks were identified on C-Cat. Absence of NiO peaks and presence of NiAl$_2$O$_4$ peaks could be observed at around 38°, 46° and 67°.

Fig. 9. (a) XRD patterns of pristine and spent catalysts, and (b) TPR profiles of pristine catalysts.

Table 3
Chloride species quantified in the syngas collected from gasifier (GS) and after catalytic tar reformer with 4 mL of F-Cat (TR).

<table>
<thead>
<tr>
<th>Chloride species</th>
<th>HCl</th>
<th>KCl</th>
<th>NaCl</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>Mean SD</td>
<td>RSD%</td>
<td>Mean SD</td>
</tr>
<tr>
<td>GS</td>
<td>0.26</td>
<td>0.15</td>
<td>58%</td>
</tr>
<tr>
<td>TR</td>
<td>0.10</td>
<td>0.06</td>
<td>60%</td>
</tr>
</tbody>
</table>

Table 4
Average yield (Nm$^3$/kg) of the main gaseous components in raw syngas from gasifier (GS), treated syngas (average of the treated syngas for different catalysts and different dosages) after catalytic tar reforming (TR) and the theoretical yield calculated based on a three steps calculation, (1) steam reforming of hydrocarbons, (2) water-gas shift reaction, (3) gasification of particulates and solid residues remained after gasification, and the thermodynamic equilibrium based on modeling.

<table>
<thead>
<tr>
<th>Main gaseous component</th>
<th>Average yield (Nm$^3$/kg)</th>
<th>Theoretical yield (Nm$^3$/kg) based on a three steps calculation and thermodynamic modeling</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td>GS Raw syngas</td>
<td>TR Treated syngas</td>
</tr>
<tr>
<td>CO</td>
<td>0.35</td>
<td>0.62</td>
</tr>
<tr>
<td>CO$_2$</td>
<td>0.30</td>
<td>0.22</td>
</tr>
<tr>
<td>CH$_4$</td>
<td>0.086</td>
<td>0.068</td>
</tr>
<tr>
<td>H$_2$</td>
<td>0.32</td>
<td>0.62</td>
</tr>
<tr>
<td>N$_2$</td>
<td>1.6</td>
<td>1.6</td>
</tr>
<tr>
<td>C$_2$-C$_5$</td>
<td>0.078</td>
<td>0.012</td>
</tr>
<tr>
<td>H$_2$O</td>
<td>0.81</td>
<td>0.62</td>
</tr>
</tbody>
</table>
which could be explained by calcination at high temperature during synthesis (i.e. 850 °C) [61]. TPR profiles, as illustrated in Fig. 9b, also suggested that in the tested catalysts, nickel was present in different states. Reduction temperatures of oxidized nickel in C-Cat between 400 and 700 °C correspond to the reduction of NiO supported on Al₂O₃. The reduction peak at 780 °C in F-Cat is attributed to the reduction of NiAl₂O₄ spinel [62], which is consistent with the results of XRD analysis. In the spent C-Cat and F-Cat, metallic nickel formed via the reduction of NiO by syngas could be observed. The formation of other crystalline phases was not observed after 2 and 20 h of operation (catalysts C-Cat and F-Cat). The XRD patterns of F-Cat after 2 h and 20 h were similar indicating high structural stability of the catalyst in raw syngas during the operation. To investigate coke formation, the fresh and spent catalysts were characterized by elemental CHN analysis (Table 5). The two spent catalysts contained considerably high carbon contents of greater than 10.8% after 2 h of tar reforming, which is likely attributed to the presence of entrained char particles from gasifier. In the spent F-Cat catalyst, carbon content decreased with an increase in operation time from 2 to 20 h indicating that no continuous carbon deposition over time occurred and accumulated char particles could be removed from the catalyst bed via either attrition or steam and CO₂ gasification reactions [35]. The presence of low concentration of sulfur species (< 100 ppmv) in the raw syngas could also suppress the carbon formation reactions on the catalyst [63]. The morphology of different areas of spent F-Cat suggests that coking of catalyst mainly occurred on the Ni particles that were not strongly attached to the alumina support. Fig. 10a illustrates the formation of carbon nanofiber on such nanoparticle. On the other hand, no visible carbon deposition was detected on Ni nanoparticles that were embedded in the alumina support (Fig. 10b). As coking of catalysts is initiated with the deposition of carbon in the Ni lattice with the subsequent transfer to the Ni particles, strong interactions between Ni and the alumina support interface, strong interactions between Ni and the alumina support could prevent the uplifting of Ni particles and the subsequent growth of carbon fiber [64].

3.3. Desulfurization of syngas

Particulates filter and dechlorination sorbents were placed before NiZn-28-HC to remove particulates and HCl. Particulates can cause fouling and clogging in the fixed bed of desulfurization sorbent. HCl reacts with ZnO forming ZnCl₂ and decreasing sulfur capacity of the fouling and clogging in the fixed bed of desulfurization sorbent. HCl reacts with ZnO forming ZnCl₂ and decreasing sulfur capacity of the sorbent. 

Desulfurization reactor was maintained at 400 °C as the higher temperature could lead to thermodynamics limitations due to increased desorption rate of sulfur species. On the other hand, lower temperature could reduce the sorption rate of COS [34]. This is particularly important for the raw syngas generated from gasification of MSW due to the presence of COS at high concentrations. For syngas desulfurization, mean concentrations of H₂S and COS species, calculated based on triitated experiments and total operation period of ~ 10 h, decreased to less than 10 ppmv in clean syngas after passing through the NiZn-28-HC sorbent (Table 2). This corresponds to more than 90% of total sulfur removal suggesting that the desulfurization of real syngas with the presence of small quantities of tar and significantly higher concentration of COS is practically feasible. The concentration of remaining sulfur compounds was below the acceptable levels for the operation of gas turbines (20 ppmv) [2], which makes NiZn-28-HC a promising sorbent for the purification of syngas at WtE plants integrated with CCGT. As the removal efficiency for H₂S was higher than 99% (Fig. S2), overall performance of desulfurization sorbent could be improved with the conversion of COS to form H₂S before desulfurization reactor. No statistically significant difference in the composition of main gases was observed before and after desulfurization reactor, indicating low catalytic activity of the sorbent at the utilized Ni-Zn loading and temperature.

XRD patterns of pristine and spent NiZn-28-HC after 2 h operation are illustrated in Fig. 11. Characteristic peaks for NiO (2θ = 38° and 43°) and ZnO (2θ = 32°, 34°, 37°) were observed indicating the presence of both oxides in Ni-Zn nanocomposite. The overall peak intensity of spent NiZn-28-HC decreased and a hump between 2θ = 25–40° appeared after desulfurization, which could be attributed to the formation of amorphous ZnS [30]. High resolution XPS spectra of NiZn-28-HC at Zn 2p, Ni 2p, S 2p and Cl 2p regions are illustrated in Fig. 12. Characteristic binding energies of 1022 and 1045 eV for Zn 2p₃/₂ and Zn 2p₁/₂ state could be attributed to Ni₂⁺ (first two peaks) and a satellite peak, respectively [22]. Highly similar XPS spectra for Zn and Ni of both pristine and spent sorbents suggested high structural consistency before and after desulfurization. Characteristic binding energies for S 2p₁/₂ and S 2p₃/₂ at 163 and 162 eV, respectively, were observed indicating the sorption of sulfur species from the syngas while the presence of S⁶⁺ peak at 168 eV implied the formation of sulfate. The formation of sulfate could be due to the reaction between Ni-Zn oxides

Fig. 10. TEM images of the spent F-Cat catalysts after 2 h operation: (a) deposition of filamentous carbon on a Ni nanoparticle loosely attached to the alumina support and (b) Ni nanoparticles embedded into alumina support with no visible carbon deposition.

and SO$_2$ or SO$_3$ species which were formed from the oxidation of H$_2$S with the Ni-Zn oxides [66]. This observation suggested that highly complex reactions could occur during the desulfurization of the real syngas with the formation of a variety of sulfur compounds such as sulfides and sulfates. Presence of Cl peaks in the region of 195–206 eV indicated that part of the Cl was adsorbed by the NiZn-28-HC which could be attributed to the incomplete HCl removal by Na$_2$CO$_3$ and/or the deposition of alkali chlorides present in syngas. This observation suggested that increased loading of Na$_2$CO$_3$ for HCl removal and/or utilization of sorbents for the removal of alkali chlorides could be beneficial for the operation of desulfurization sorbent. FESEM micrographs as showed in Fig. 13 illustrate the surface morphology of pristine and spent NiZn-28-HC. Based on the similar features, the structure can of sorbent was not affected by the desulfurization of syngas. This could be attributed to the presence of Ni dopant which acted as a promoter to stabilize the lattice expansion and contraction effects [67]. Results of the characterization of the spent NiZn-28-HC showed high similarity with the results from a bench-scaled desulfurization set-up and suggested that the spent sorbent could be regenerated and reused as demonstrated previously [28].

In average, the amount of S captured by the NiZn-28-HC was estimated at 2.0 ± 0.8 mg/g of HC or equivalently, 113 ± 46 mg/g of Ni-Zn oxides. This is statistically comparable to the amount of S captured

---

**Fig. 11.** XRD patterns of pristine and spent Ni-Zn-28-HC.

**Fig. 12.** High resolution XPS spectra at (a) Zn 2p, (b) Ni 2p, (c) S 2p and (d) Cl 2p for pristine and spent NiZn-28-HC.
The slight reduction on the capturing of Scould be attributed to the strong interactions between Ni particles and alumina support. Considering the tar dew point and specific concentration limits for different tar compounds, the treated syngas could be applied in the gas engines/turbines and SOFCs. Regenerable NiZn-28-HC desulfurization sorbent demonstrated excellent performance for syngas produced from MSW gasification, which consisted of significantly higher proportion of COS than H2S and only traces of COS (7 ppmv). In addition, the incomplete HCl removal by Na2CO3 and/or the deposition of alkali chlorides present in real syngas could reduce the S removal capacity and efficiency.

4. Conclusions

A hot syngas purification system with downward cascading of syngas temperature was developed with satisfactory performance to treat raw syngas generated from downdraft gasification of real MSW and to produce clean syngas potentially applicable in more efficient downstream applications such as gas engines, CCGTs, SOFCs and hybrid energy recovery systems. Up to 90% removal efficiency of tar and sulfur compounds were achieved. Operational stability was demonstrated for the fluidized-bed catalytic tar reformer with nano-structured F-Cat treating dirty raw syngas directly from gasifier without any preliminary treatment. The strong resistance of catalyst to coking and poisoning could be attributed to the strong interactions between Ni particles and alumina support. Considering the tar dew point and specific concentration limits for different tar compounds, the treated syngas could be applied in the gas engines/turbines and SOFCs. Regenerable NiZn-28-HC desulfurization sorbent demonstrated excellent performance for syngas produced from MSW gasification, which consisted of significantly higher proportion of COS than H2S and trace amounts of tar and HCl. Multiple impurities in raw and treated syngas were sampled and analyzed at different sampling points throughout this purification system and provided information on the evolution and distributions of these compounds. Comparison of the composition of raw and treated syngas with that at thermodynamic equilibrium provided insight into the differences between experimental and thermodynamic modeling results, particularly, on the kinetic constraints of steam reforming of hydrocarbons which include tar, CH4 and C2-C5 gases, water-gas shift reaction and gasification of particulates in syngas and solid residues remained in the gasifier. These constraints could be caused by the poisoning of catalysts by the impurities (HCl and sulfur species) and the slow reaction kinetics between char, H2O and CO2. Complementary and comparable results for tar contents were obtained by using two different tar sampling methods, namely collected tar and SPA tar, indicating the applicability of both methods. Strong correlations were formed for these two methods to perform the correction of the loss of volatile tar compounds when the SPA sampling was carried out for a stream of syngas at higher temperature (~ 850 °C) and reduced pressure (~ 95 kPa), and in this case, toluene from SPA tar.

Acknowledgements

This research is supported by the National Research Foundation, Prime Minister’s Office, Singapore and the National Environment Agency, Ministry of the Environment and Water Resources, Singapore, under the Waste–to–Energy Competitive Research Programme (WTE CRP 1501 105). The authors also acknowledge the management of Nanyang Environment and Water Research Institute and Economic Development Board, Singapore for the support.

Appendix A. Supplementary material

Supplementary data to this article can be found online at https://doi.org/10.1016/j.apenergy.2019.01.031.

References:


