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Dye-sensitized Pt@TiO₂ core–shell nanostructures for the efficient photocatalytic generation of hydrogen

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Full Research Paper

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Keywords:
charge transfer; dye-sensitization; photocatalysis; photocatalyst; solar fuels; water splitting

Abstract
Pt@TiO₂ core–shell nanostructures were prepared through a hydrothermal method. The dye-sensitization of these Pt@TiO₂ core–shell structures allows for a high photocatalytic activity for the generation of hydrogen from proton reduction under visible-light irradiation. When the dyes and TiO₂ were co-excited through the combination of two irradiation beams with different wavelengths, a synergic effect was observed, which led to a greatly enhanced H₂ generation yield. This is attributed to the rational spatial distribution of the three components (dye, TiO₂, Pt), and the vectored transport of photogenerated electrons from the dye to the Pt particles via the TiO₂ particle bridge.

Introduction
Since Honda and Fujishima reported the effective hydrogen evolution from water splitting by a TiO₂ and Pt electrode in a photoelectrochemical cell in the early 1970s [1], TiO₂ has received extensive attention as one of the promising semiconductor photocatalysts, because of its high chemical stability, low cost and non-toxicity [2-5]. However, it suffers from the wide band gap (3.2–3.4 eV), which restricts the utilization of visible light, and the high recombination rate of photogenerated electrons and holes often leads to low quantum yields and a poor photocatalytic activity [6]. Tremendous efforts have been made to improve the photocatalytic performance of TiO₂. One typical strategy is prolonging the lifetime of the electron–hole pair through deposition of noble metal (e.g., Pt) nanoparticles as co-catalysts that can act as electron-sinks to achieve effective charge separation on TiO₂ [7-11]. Dye-sensitization has been widely used to enable visible light harvesting by wide band gap semiconductors. Since the seminal work reported by O’Regan and Grätzel in 1991 [12], various types of dyes have been explored, and some of them allow for the reduction of protons into hydrogen gas through visible-light-driven photocatalytic processes [13-17]. Herein, we use erythrosin B (ErB) sensitized Pt@TiO₂ core–shell nanoparticles for the highly-efficient...
photocatalytic generation of hydrogen under visible-light irradiation. In this rational design of the ternary structure, the TiO$_2$ particle acts as a bridge that allows for the effective electron transfer pathway of excited ErB$\rightarrow$TiO$_2$$\rightarrow$Pt. Importantly, we found that when the TiO$_2$ bridges are excited simultaneously, the dye-sensitization-driven H$_2$ evolution showed a much higher efficiency as compared to the situation with no excitation of TiO$_2$. This kind of synergic effect reveals a new direction for improving the efficiency of composite photocatalysts by using selective excitation wavelengths.

Results and Discussion

The Pt@TiO$_2$ core–shell nanoparticles were prepared through a hydrothermal process by using Pt nanoparticles and TiF$_4$ as the precursor. The crystalline structure was determined by XRD, as shown in Figure 1. After the hydrothermal reaction, the TiO$_2$ was transformed into anatase phase, which could be well indexed to the standard anatase TiO$_2$ (JCPDS Card No. 83-2243). The three additional diffraction peaks shown in Figure 1 could be assigned to the face-centered metallic Pt phase, with the positions at 40.0°, 46.6° and 67.9° representing the spacing of the (111), (200) and (220) planes, respectively. This indicates the retaining of the Pt nanoparticle cores after the hydrothermal reaction. As a control sample, Pt/TiO$_2$ was prepared through the photodeposition of Pt (1% in mole fraction) onto pure TiO$_2$ particles that were prepared through the same hydrothermal method without using Pt nanoparticles. Due to the low Pt loading, we could only observe a weak diffraction peak at 40° corresponding to the Pt (111) lattice spacing, which could indicate successful loading of metallic Pt nanoparticles onto TiO$_2$ particles. The molar ratio of Pt to Ti was estimated to be about 6.7% according to EDX analysis (Figure S1, Supporting Information File 1).

Figure 1: XRD patterns of Pt@TiO$_2$ and Pt/TiO$_2$ samples.

Figure 2: TEM and SEM images of the Pt@TiO$_2$ sample. (A) (B) TEM images of Pt@TiO$_2$. (C) HRTEM images of Pt@TiO$_2$. (D) SEM image of Pt@TiO$_2$.

Figure 3 shows the UV–vis diffuse reflectance spectra of the Pt@TiO$_2$ core–shell nanostructures and the Pt/TiO$_2$ control sample. The absorption from 250 to ca. 380 nm can be attributed to the band edge absorption of anatase TiO$_2$. The band gaps of both samples are calculated according to the modified...
Kubelka–Munk function $\left[ F(R_g)E \right]^{1/2}$ vs the energy of absorbed light, $E$ [18]. And the plots shown in the inset of Figure 3 reveal the band gap value as 3.3 eV for Pt@TiO$_2$ and 3.2 eV for Pt/TiO$_2$, which indicates that the combination of Pt nanoparticles with TiO$_2$ did not significantly influence the band gap energy of the TiO$_2$ component.

The photocatalytic tests were carried out by suspending the Pt@TiO$_2$ sample (5 mg) in an aqueous solution containing 0.2 wt % ErB and 15 wt % triethanolamine (TEOA) that acts as the electron donor. In this study, we employed two different irradiation wavelengths for the purpose of separating the excitation of TiO$_2$ and ErB, which will help us to explore the potential synergic effect between the two excitations. The primary irradiation with a wavelength of 550 ± 20 nm (light A) is to excite ErB since the main absorbance peak of ErB is located at about 550 nm. A secondary irradiation with a wavelength of 400 ± 10 nm (light B) is to excite the defect/impurity states of TiO$_2$, while ErB exhibit a minimum absorption in this range.

As shown in Figure 4, after the individual irradiation with light A (550 ± 20 nm) or light B (400 ± 10 nm) for 2 h, the ErB-sensitized Pt@TiO$_2$ core–shell structure showed generated H$_2$ amounts of 4.5 μmol and 5.3 μmol, respectively. However, when light A and light B are used simultaneously, to our surprise, the 2 h irradiation led to a H$_2$ amount of 15.9 μmol, which is significantly higher than the sum of the two generated H$_2$ amounts under individual irradiation of light A and B (9.8 μmol). This observation suggests that in the ErB-sensitized Pt@TiO$_2$ core–shell structure, a synergic effect exists between the excitation of ErB and TiO$_2$, which plays an important role in the photocatalytic hydrogen generation.

The observed synergic effect could be attributed to the electron transport in TiO$_2$ particles. Since the dye-sensitization induces an electron transfer from the excited ErB to TiO$_2$, these electrons have to be transported through the TiO$_2$ particles with a maximum distance of ca. 60 nm to reach the Pt surface for the reduction of protons to H$_2$. When TiO$_2$ is simultaneously excited by the 400 ± 10 nm light, though it is weak, the charge carrier concentration in TiO$_2$ becomes higher, which increases the conductivity of TiO$_2$. Thereby, the vectored electron transfer from ErB to the core Pt particle via TiO$_2$ bridges becomes more effective, which leading to enhanced yield of H$_2$. The principle is depicted in Figure 5, and the energy diagram is shown in Figure S3 (Supporting Information File 1).
and 1.9 μmol, respectively. While the simultaneous irradiation with light A and B for 2 h led to the generation of 7.9 μmol H₂, which is only slightly higher than the sum of the H₂ generation amount under individual irradiation of light A and B (7.2 μmol). This indicates that the synergistic effect may also exist in the Pt-loaded TiO₂ particles, but with a less significant role as compared to the Pt@TiO₂ core–shell structure. This may be because the post-loaded Pt nanoparticles are randomly distributed on the outer surfaces of TiO₂ particles, thus the electron transfer path from ErB to Pt becomes less oriented through TiO₂. In addition, the Pt nanoparticles would occupy some reactive sites on the TiO₂ surface, which also reduces the efficiency of the dye-sensitization.

Conclusion

We have prepared Pt@TiO₂ core–shell nanostructures through a one-step hydrothermal method. Upon ErB sensitization, the Pt@TiO₂ core–shell photocatalysts exhibit high visible-light activity for the generation of H₂ from proton reduction. Significantly, we observed a synergistic effect that allows for a greatly enhanced activity for the H₂ generation when the ErB and TiO₂ are co-excited through the combination of two irradiation beams at different wavelengths. The enhancement is attributed to the rational spatial distribution of three components (ErB, TiO₂, Pt), and the vectored transport of photogenerated electrons from ErB to Pt particles via the TiO₂ particle bridge. The presented core–shell structures and the observed synergistic effect would provide a new direction for improving the efficiency of composite photocatalysts by using selective excitation wavelengths.

Experimental

Seed growth of 30 nm Pt nanoparticles: Seed Pt nanoparticles were prepared first. Typically, an aqueous solution of H₂PtCl₆ (3.8 mM, 7 mL) was added into 90 mL deionized (DI) water and heated to boil under stirring. After that, 2.2 mL aqueous solution containing 1% trisodium citrate and 0.05% citrate acid was added, followed by a quick injection of freshly prepared NaBH₄ solution (0.08%, 1.1 mL). After 10 min, the solution was cooled down to room temperature, and used as the solution of seed nanoparticles (≈4 nm). For further growth into 16 nm Pt nanoparticles, 1 mL of this Pt seed solution was added into 29 mL DI water, followed by the addition of 0.09 mL solution of H₂PtCl₆ (0.2 M) and 0.5 mL of a solution containing 1% sodium citrate and 1.25% L-ascorbic acid. The solution was kept under stirring and heated to boil. After 30 min, the solution was cooled down to room temperature and used as the 16 nm Pt seed solution for further growth into 30 nm Pt nanoparticles. In a typical run, 4 mL of the 16 nm Pt particle solution was mixed with 26 mL DI water. Then 0.09 mL solution of H₂PtCl₆ (0.2 M) was added, followed by addition of 0.5 mL solution containing 1% trisodium citrate and 1.25% L-ascorbic acid. The solution was kept stirring and heated to boil, and after 30 min of boiling, the solution was cooled down to room temperature and used the solution of 30 nm Pt nanoparticles.

Preparation of Pt@TiO₂ core–shell nanostructures: The Pt@TiO₂ core–shell nanostructures were synthesized through a hydrothermal method. Typically, 15 mL solution of the as-prepared 30 nm Pt nanoparticles was mixed with 4.5 mL aqueous solution of TiF₄ (0.04 M). The mixture was kept stirring for 10 min, diluted into 80 mL DI water, and then transferred into a 100 mL Teflon-lined stainless steel autoclave, which was treated at 180 °C for 24 h. After that, the product was cooled down to room temperature, centrifuged and washed with deionized water for three times, and dried in a vacuum oven.

Characterizations: The crystalline phases of the samples were examined by powder X-ray diffraction (XRD) on a Shimazu XRD-6000 X-ray diffractometer (Cu Kα radiation) with a scanning speed of 2°/min in the 2θ range from 20 to 80°. Diffuse reflectance UV–vis spectra were acquired on a Lambda 750 UV–vis–NIR spectrophotometer (Perkin Elmer, USA). The morphology of the Pt@TiO₂ nanocomposites were investigated by field emission scanning electron microscopy (SEM, JEOL JSM-7600F) with energy-dispersive X-ray analysis system and transmission electron microscopy (TEM, JEOL JEM-2100) at an accelerating voltage at 200 kV.

Photocatalytic generation of H₂ with erythrosin B-sensitized Pt@TiO₂ core–shell particles: In a typical run, 5 mg Pt@TiO₂ photocatalyst was dispersed into 10 mL of an aqueous solution containing triethanolamine (TEOA, 15 wt %) as electron donor and erythrosin B (0.2 wt %) as the photo-sensitizing dye. The suspension was sealed in a quartz vessel and purged with Argon for 30 min to remove the residual oxygen. After that, the vessel was exposed under a 300 W Xenon lamp (MAX-302, Asahi Spectra Co. Ltd.) coupled with a band pass filter (λ = 400 ± 10 nm or 550 ± 20 nm) to evaluate the photocatalytic H₂ generation yield. The gas products were analyzed periodically by an Agilent 7890A gas chromatograph (GC) with a thermal conductivity detector (TCD).

Supporting Information

Supporting Information File 1
Additional experimental data.
[http://www.beilstein-journals.org/bjnano/content/supplementary/2190-4286-5-41-S1.pdf]
Acknowledgements
This work was financially supported by NTU seed funding for Solar Fuels Laboratory, MOE AcRF-Tier1 RG 44/11, MOE AcRF-Tier2 (MOE2012-T2-2-041, ARC 5/13), and CRP program (NRF-CRP5-2009-04) from the Singapore National Research Foundation.

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The definitive version of this article is the electronic one which can be found at: doi:10.3762/bjnano.5.41